

App MEL

Chemistry APP

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APP

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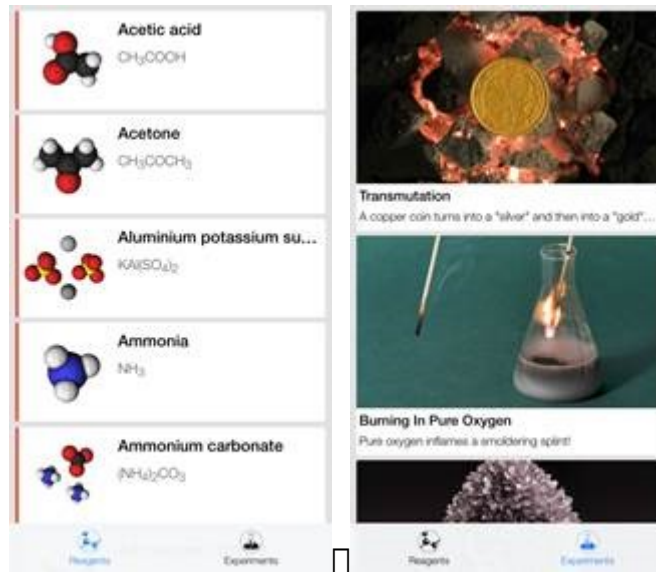
### n MEL Chemistry APP

MEL Chemisty app MEL Science how to pass the knowledge about science in an easy, interesting and effective way MEL Science 2017 SVOD (Silicon Valley Open Doors) iOS android MEL Chemistry 1



1 MEL Chemistry android ios

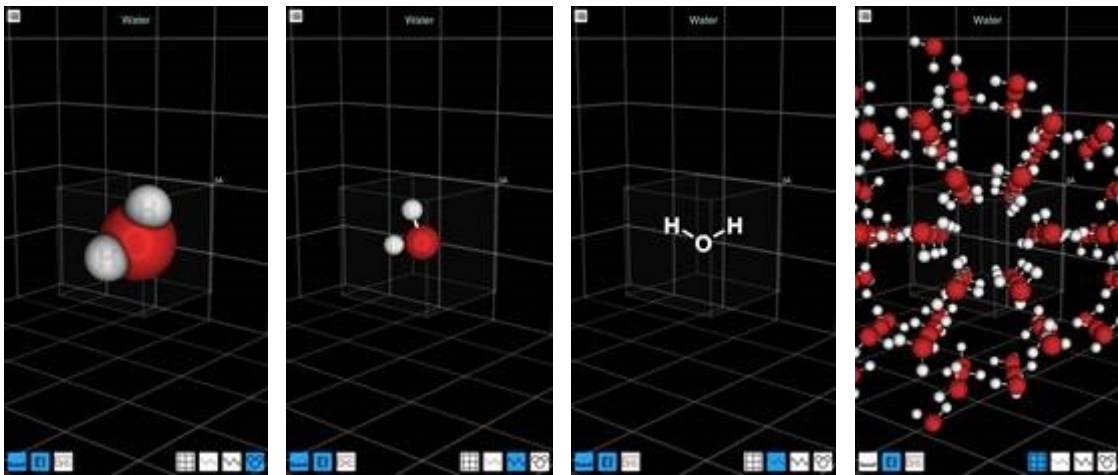
MEL Chemistry Reagents Experiments 2  
 Reagents  
 Experiments MEL Chemistry



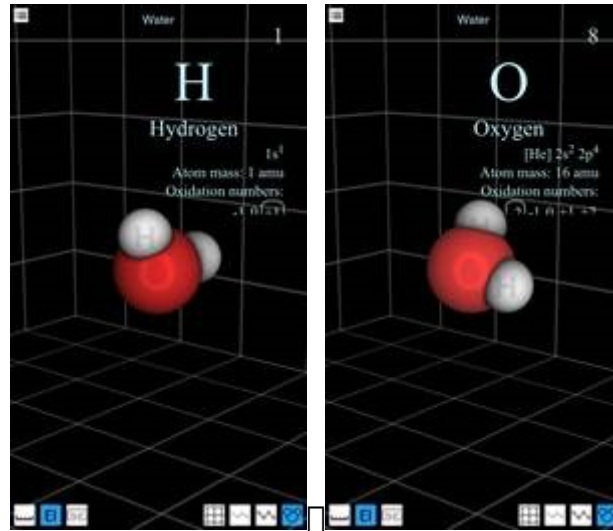
2 MEL Chemistry

Reagents

Reagents 80  
 7 3  
 3

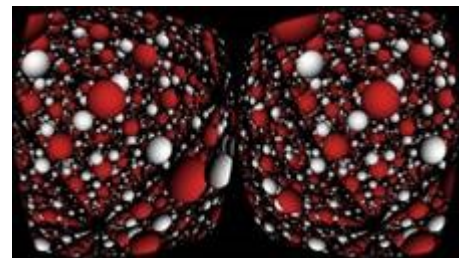
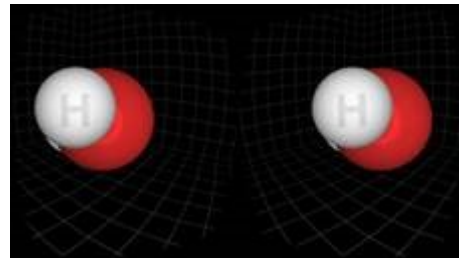


3 Reagents



4

SE AI
  
 Google Cardboard<sup>4</sup> app



5 Google Cardboard VR VR

### Experiments


Experiments

MEL Chemistry ( )
   
 Difficulty Danger Duration
   
 Youtube
   
 (reagents)

## Burning magnesium

Magnesium burns in air with a very bright flame

Difficulty: ■ ■ ■ ■ ■ Danger: ■ ■ ■ ■ ■ Duration: 7 minutes



$$2\text{Mg} + \text{O}_2 \xrightarrow{\text{T}^\circ} 2\text{MgO}$$

**Reagents** →

- Magnesium

**Safety**

- Conduct the experiment on the tray.
- Take protective gloves off before lighting the candle.
- Handle magnesium only with clean and dry hands.

MEL Science

MEL Chemistry sets reagents


Magnesium

## Magnesium

Molecular formula: Mg  
 Chemical compound: magnesium  
 Common name Other names: [Wikipedia page](#)  
[MSDS \(material safety data sheet\)](#)

### 6 Experiments

MEL Chemistry (step-by-step instruction) (expected result) (MEL Chemistry)



**Expected result**

Magnesium burns in air very actively, causing a bright glow and releasing a generous amount of energy. The main product of magnesium with oxygen reaction is magnesium oxide MgO.

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Magnesium burns in air very actively, causing a bright glow and releasing a generous amount of energy. The main product of magnesium with oxygen reaction is magnesium oxide MgO.

**Disposal**

Dispose of the experiment residues along with regular household trash.

**Scientific description**

**Why does magnesium burn so brightly?**

Magnesium is a very active metal. When reacting with oxygen from

**That's interesting!**

**What is magnesium used for?**

Magnesium burns extremely brightly, and this property found its use. The spectrum of light released during magnesium burning has a significant ultraviolet component. That's why it was used in photography for a while. Mixtures of magnesium with various oxidizers (barium nitrate Ba(NO<sub>3</sub>)<sub>2</sub>, potassium chlorate KClO<sub>3</sub> or potassium permanganate KMnO<sub>4</sub>) were used as a photographic flash because photoplates back then were very sensitive to ultraviolet.

Currently, metallic magnesium is used in signal and illumination flares, fireworks, flash grenades, and tracer bullets as a bright white light source.

### 7 Experiments

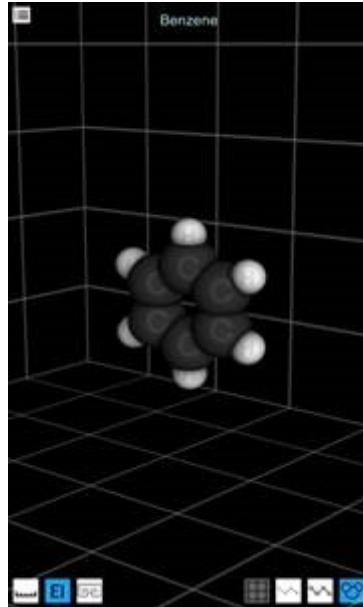
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MEL Chemistry ( )

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MEL Chemistry NaCl Na<sup>+</sup> 6

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MEL Chemistry APP Reagent  
80  $\text{SiO}_2$   $\text{CuO}$   $\text{CaO}$   
Experiments

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